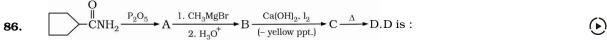


Date Planned ://	Daily Tutorial Sheet-7	Expected Duration : 90 Min
Actual Date of Attempt : / /	Level-2	Exact Duration :



- (в) —соон

(c) _______

- (D)
- **87.** Which of the following reactions is not reversible?
 - (A) $RCOOR' + H_2O \xrightarrow{H^+} RCOOH + R'OH$ (B)

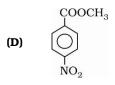
 $RCOOR' + NaOH \longrightarrow RCOONa + R'OH$

- (C) $RCOOH + R'OH \longrightarrow RCOOR' + H_2O$
- $RCH = CH_2 \xrightarrow{H_2O/H^+} RCH(OH)CH_3$
- *88. Which of the following will loose CO_2 on heating:
 - (A) methyl malonic acid
- **(B)** succinic acid
- (C) 2, 2-dimethyl acetoacetic acid
- (D) Lactic acid
- ***89.** The ester, RCOOR' can be prepared by :
 - (A) esterification of RCOOH with R'OH
 - **(B)** esterification of (RCO)₂O
 - (C) Baeyer-Villiger oxidation of RCOR' with peroxybenzoic acid
 - (D) reaction of RCOCl with R'OH
- ***90.** Which of the following statement is (are) true?
 - (A) At room temperature, formyl chloride is present in the form of CO and HCl
 - **(B)** Acetamide behaves as a weak base as well as a weak acid
 - (C) $CH_3CONH_2 \xrightarrow{LiAlH_4} C_2H_5NHCH_3$
 - (D) $CH_3CONHR \xrightarrow{H_3O^+} CH_3COOH + RNH_2$
- **91.** Which of the following will undergo alkaline hydrolysis most rapidly?





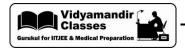




PARAGRAPH FOR QUESTIONS 92 - 93

An organic compound X, $C_4H_8O_3$ evolves CO_2 on reaction with NaHCO $_3$. It undergoes dehydration to give an unsaturated compound Y which is highly stable. Y on oxidation with hot acidic $KMnO_4$ gives P and Q.

- **92.** X is:
 - (A) CH₃CH₂CH(OH)COOH and is optically active
 - **(B)** CH₃CH(OH)CH₂COOH and is optically inactive
 - (C) CH₃CH(OH)CH₂COOH and is optically active
 - (D) HOCH₂CH₂CH₂COOH and is optically inactive



- **93.** P and Q are respectively:
 - **(A)** Both are $(COOH)_2$

- **(B)** Both are CH_3COOH
- (C) CH₃COOH and (COOH)₂
- **(D)** CH_3CHO and $CHO \cdot COOH$
- 94. HOOC—P. The barium salt of P on heating will give :

 COOH
 CH₂COOH
 - (A) HOOC- =0 CH₂OOH

(c)

- **(D)** CH:
- **95.** Consider the following sequence of reactions.
 - Ph H_3C = N H_2SO_4 A $\frac{1. H_2SO_4, heat}{2. H_3O^+}$ C + D

The products (B) and (C) are:

- (A) PhCOOH and CH_3NH_2
- **(B)** $CH_3COOH \text{ and } PhNH_2$
- (C) $PhCOCH_3$ and NH_2OH
- **(D)** PhCONHCH₃ and CH₃CONHPh

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